

Prior Learning					
Lesson Number		AQA Spec	Title	Content	Assessment/ Homework
Original	WSFG				
1 and 2	1 and 2	4.3.1.1	Conservation of mass and balanced equations	<ul style="list-style-type: none"> Explore ideas about the conservation of mass. Consider what the numbers in equations stand for. Write balanced symbol equations. 	
3	2	4.3.1.2	Relative formula mass	<ul style="list-style-type: none"> Identify the relative atomic mass of an element Use relative atomic masses to calculate relative formula masses. 	
4	3	4.3.1.3	Mass changes when gases are in reactions	<ul style="list-style-type: none"> Find out how mass can be gained or lost during a reaction Find the mass of carbon dioxide released per gram of copper carbonate decomposed Interpret data and graph skills practice Maths skills: change the subject of an equation Carry out multistep equations 	

6	5	4.3.2.1	Moles	<ul style="list-style-type: none"> Describe what is meant by the term moles. Avogadro's constant Calculate number of moles Calculate the mass of a given number of moles of a substance 	Assessment (15 marks)
7	6	4.3.2.2 Higher tier only	Amounts of substances in equations	<ul style="list-style-type: none"> Predict the mass of reactant or product that will be produced from a given balanced equation 	
8	7	4.3.2.3 4.3.2.4 Higher tier only	Using moles to balance equations	<ul style="list-style-type: none"> Convert masses in grams to amounts in moles Use moles to balance equations 	Assessment 2: Online Quiz
9	8	4.3.2.5 Higher tier only	Concentration of solutions	<ul style="list-style-type: none"> Calculate the concentration of substances using mass of solute Re-arrange equation to calculate mass from the concentration Calculate concentration from moles 	
End of chapter test					
Where we will use these ideas again					

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10 TRIPLE	9	4.3.4	Using concentrations of solutions (Titrations)	<ul style="list-style-type: none"> Describe how to carry out titrations. Calculate concentrations in titrations in mol/dm³ and in g/dm³. Explain how the concentration of a solution in mol/dm³ is related to the mass of the solute and the volume of the solution. 	
11 TRIPLE	10	4.4.25	Required practical: Finding the reacting volumes of acid and alkali by titration	<ul style="list-style-type: none"> Use an acid to neutralise a known volume of alkali. Use a burette to determine the volume of an acid needed. Use the results to determine the concentration of an alkali. 	
12 TRIPLE	11	4.3.3.1	Key concept: Percentage yield	<ul style="list-style-type: none"> Calculate the percentage yield from the actual yield. Identify the balanced equation needed for calculating yields. Calculate theoretical product amounts from reactant amounts. 	

13 TRIPLE	12	4.3.3.2	Atom economy	<ul style="list-style-type: none"> • Identify the balanced equation of a reaction. • Calculate the atom economy of a reaction to form a product. • Explain why a particular reaction pathway is chosen. 	
14 TRIPLE	13	4.3.5	Amounts of substance in volumes of gases	<ul style="list-style-type: none"> • Explain that the same amount of any gas occupies the same volume at room temperature and pressure (rtp). • Calculate the volume of a gas at rtp from its mass and relative formula mass. • Calculate the volumes of gases from a balanced equation and a given volume of a reactant or product. 	
End of chapter test					
Where we will use these ideas again					